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Carbohydrate

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Part I - Organic  
Compounds and  
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Chapter 22

(Organic

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Reduction

Reactions Nature

of Oxidation and

Reduction

Questions 1.

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Explain in your own words the following oxidation processes: a. Gain of oxygen b. Loss of hydrogen c. Loss of electrons 2.

Explain in your own words the following reduction processes: a. Loss of oxygen b.

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13.3 Liquids and Solids In your textbook, read about liquids and solids. In the space at the left, write true if the statement is true; if the statement is false,

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704 CHAPTER 22

Identify the product that balances the

following nuclear

reaction:  ${}_{84}^{212}\text{Po}$

$\rightarrow {}_2^4\text{He} + 1$ . The

total mass number and atomic number must be equal on

both sides of the

equation.  ${}_{84}^{212}\text{Po}$

$\rightarrow {}_2^4\text{He} +$  mass

number:  $212 - 4$

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$Z=82$  atomic  
number:  $84 - 2$   
 $= 82$ . The nuclide  
has a mass number  
of 208 and an  
atomic number of  
82, 208  $^{82}\text{Pb}$ . 3.

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Structural Formulas  
of Alkanes □ SG

22.1 & 22.2

Structures of

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Alkenes & Alkynes

SG 22.3 & 22.4 SG

23.1, 23.2, 23.3

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stability of a  
nucleus Define and  
relate the terms  
decay series,  
parent nuclide ...

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Problems 13. Use  
dimensional  
analysis and the  
given densities to  
make the following  
conversions. a.  
14.8 g of boron (B)  
to cubic  
centimeters of  
boron. The density  
of boron is 2.34 g/

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$\text{cm}^3 \cdot (14.8 \text{ g/cm}^3) \cdot (1.32 \cdot 10^3 \text{ cm}^3) = 6.32 \cdot 10^3 \text{ g}$

2.8 L of argon (Ar)  
to grams of argon.

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1) No, it is not a proper answer; you do not know whether the professor meant

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number 20 or 20  
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problems. Twenty  
is a quantity and  
does not make  
sense without a

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unit. 3) Middle  
Zeros: Zeros that  
appear between  
other nonzero  
digits are ALWAYS  
significant.

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