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Modern Chemistry Section 3 Gases

Modern Chemistry 97 Gases CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. _____ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2.

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Three of the primary components of air are carbon dioxide, nitrogen, and oxygen. In a sample containing a mixture of only these gases at exactly 1 atm, the partial pressures of carbon dioxide and nitrogen are given as $P_{CO_2} = 0.285$ torr and $P_{N_2} = 593.525$ torr. What is the partial pressure of oxygen?

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SECTION 1 continued 5. Convert a pressure of 0.200 atm to the following units: 152 a. mm Hg 20.3 b. kPa 6. When an explosive like TNT is detonated, a mixture of gases at high temperature is created. Suppose that gas X has a pressure of 50 atm, gas Y has a pressure of 20 atm, and gas Z has a pressure of 10 atm. 80 atm a. What is the total ...

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Chapter 11 Review Gases Section 2 Answers Modern Chemistry

Chapter 11 Review Gases Section 1 Answers Modern Chemistry 97 Gases CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. _____ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided ...

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Modern Chemistry 76 Quiz Section Quiz: Gas Volumes and the Ideal Gas Law In the space provided, write the letter of the term or phrase that best completes each sentence or best answers each question. _____ 1. At the same temperature and pressure, balloons of equal volume always contain

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